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Concept 1: CHNOPS: The Six Most Abundant Elements of Life

1. By what bond do most biomolecules form? List the 6 most abundant elements found in biomolecules.

Concept 2: Valence and Covalent Bonding

2. What is the valence of each of these elements? Draw their stick models.

Concept 3: Organic Molecules: Hydrocarbons

3. What is special about Carbon?
4. Name the element studied in Organic Chemistry.
5. Describe what a hydrocarbon is, give an example, draw its stick model.

Concept 4: Isomers

6. Describe an Isomer, give an example.
7. Name and give the differences between the 3 types of isomers.

Concept 5: Polarity

8. How is polarity and covalent bonding related?
9. What is electronegativity?
10. Draw stick molecule of water showing it's charges.

Concept 6: The Functional Groups

11. Fill out the following chart for each of the functional groups:

Group name	Symbol	Description	Example	Other info

Complete self quiz: Put answers below:

1. _____ 2. _____ 3. _____ 4. _____ 5. _____ 6. _____ 7. _____ 8. _____

Key

Ions

How are ions made from neutral atoms?

Why?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Model 1 to complete the following table.

	Metal or Nonmetal	Is the number of protons the same in the atom and the ion?	Is the number of neutrons the same in the atom and the ion?	Is the number of electrons the same in the atom and the ion?	Charge on the ion
Lithium	metal	yes	yes	no	1+
Magnesium	metal	yes	yes	no	2+
Aluminium	metal	yes	yes	no	3+
Fluorine	nonmetal	yes	yes	no	1-
Oxygen	nonmetal	yes	yes	no	2-
Nitrogen	nonmetal	yes	yes	no	3-

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

number of electrons

3. Examine the isotope symbols in Model 1.

a. Where is the ion charge located in the isotope symbol?

to the right of the chemical symbol

b. Is a charge indicated on the neutral atoms? If yes, where is it located?

NO.

4. Which subatomic particle carries a positive charge?

proton

5. Which subatomic particle carries a negative charge?

electron

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Model 1.

protons - electrons Lithium ion: $3 - 4 = -1$

or Nitrogen ion: $7 - 10 = -3$

(protons) + (- electrons)

