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## What is the difference between thermal energy and heat energy

The energy measured by the temperature technique radiation in the visible light can be seen in this hot metals. The term "thermal energy" is used differently, and often loosely in different contexts. Refers to several distinct physical concepts, such as internal energy, or as enthalpy, from a body of matter and radiation; or as heat, defined as a type of energy transfer (as is thermodynamic work); or as the characteristic energy of a degree of freedom,  $k_B t$ , in a system described in terms of its particulate microscopic constituents, wherein  $t$  denotes temperature  $k_B$  denotes the boltzmann constant. Relationship with heat and internal energy in thermodynamic, the heat is energy in transfer to or from a thermodynamic system, by mechanisms that do not the thermodyic work or transfer of matte. [1] [2] [3] The heat refers to an amount transferred between the systems, not to a property of any system, or 'contained' within it. [4] On the other hand, internal energy and enthalpy are properties of a single system. The heat and work depend on the way in which an energy transfer occurred, while the internal energy is a property of the state of a system and can be understood without knowing how the energy arrived there. In a stating mechanical account of an ideal gas, in which the molems move independently between instant collages, the internal energy is the total sum of the cinemic energies of the independent particles of the  $G^A$  © and this cycle movement is the source and the effect of heat transfer through the limit of a system. For a san that do not have particulate interactions, except for instant collages, the term "thermal energy" is effectively synonymous with "internal energy". In many texts of statum fans, "Temica Energy" refers to  $kt$ , the product of constant and the absolute temperature of Boltzmann, also written as  $k_B t$ . [5] [6] [7] [8] [9] in a material, especially in condensed matte, such as a liquid or a solid, in which the constituent particles, such as moleps or nonsons, interact strong with Others, the energies of these interactions strongly contribute to the internal energy of the body, but are not simply apparent in temperature. The term 'thermal energy' is also applied to the energy carried by a heat flow, [10] although this can also be called heat or amount of heat. Historic context in a lecture of 1847 titled "In Matência, Living and Heat Forces," James Prescott Joule characterized several terms that are closely related to ethermic and heat energy. He identified the terms latent heat and sensitive heat as heat forms that affect distinguished physical phenomena, namely the potential and kinemic energy of the particles, respectively. [11] He described latent energy as the energy of the interaction in a particular particle configuration, this is a potential form of energy and sensible heat as an energy affecting the measured temperature Termometer due to the thermal energy, which he called a living force. Thermal Energy Inactile If the minimum temperature of the environment of a system is  $t$  and the system entropy is  $s$ , a part of the internal energy of the system

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